

surface is not carbon rich, decomposition of SiC occurs and titanium silicides and titanium carbide are formed.

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**Registry No.** Ti, 7440-32-6; B, 7440-42-8; SiC, 409-21-2; TiC, 12070-08-5; SiO<sub>2</sub>, 7631-86-9; TiO<sub>2</sub>, 13463-67-7; titanium silicides, 12738-91-9.

## Additions and Corrections

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1989, Volume 1.

**Duncan W. Bruce,\* David A. Dunmur, Peter M. Maitlis, Peter Styring, Miguel A. Esteruelas, Luis A. Oro, M. Blanca Ros, José-Luis Serrano, and Eduardo Sola:** Nematic Phases in Ionic Melts: Mesogenic Ionic Complexes of Silver(I).

Page 479. In the paper, we reported the phase behavior of [Ag(4-OPhVPy)<sub>2</sub>][DOS] incorrectly. The correct behavior is described according to

K-S <sub>A</sub>	147 °C	$\Delta H = 32.68 \text{ J g}^{-1}$	$\Delta S_m/R = 8.1$
S <sub>A</sub> -N	166 °C	$\Delta H = 0.7 \text{ J g}^{-1}$	$\Delta S_m/R = 0.6$
N-I	178 °C	$\Delta H = 0.2 \text{ J g}^{-1}$	$\Delta S_m/R = 0.07$

Further, the mesophase does not change on treatment of the sample with wet acetone. Simply the strongly homeotropic texture of the S<sub>A</sub> phase is replaced by the more readily identifiable focal conic texture.

1991, Volume 3.

**David K. Liu,\* Roger J. Chin, and Angela L. Lai:** Photochemical Vapor Deposition of Iron-Cobalt Thin Films: Wavelength and Temperature Control of Film Compositions.

Page 14. In the second paragraph, sixth sentence, the units for saturation moment should be  $\mu_B/\text{mol}$ .